

ADVANCED CHEMICAL-MECHANICAL DEWATERING OF FINE PARTICLES

By

Ramazan Asmatulu

Dissertation submitted to the Faculty of the
Virginia Polytechnic Institute and State University
In partial fulfillment of the requirement for the degree of

Doctor of Philosophy

In

Material Science and Engineering

Dr. R.-H. Yoon, Chairman

Dr. G.H. Luttrell

Dr. G.T. Adel

Dr. J.P. Wightman

Dr. W.T. Reynolds

March 12, 2001

Blacksburg, Virginia

Keywords: Dewatering, Fine Particles, Moisture, Filtration, Centrifugation, Flotation, Coagulation, Oxidation, Buchner Filter, Drum, Disc, HBF, Surfactant, Polymer, Electrolyte, Coal, Chalcopyrite, Sphalerite, Galena, Clay, Phosphate, PCC, Silica, Talc, XPS, PZC, BET, AFM, Hydrophobic Force, Hydrophobicity, Hydrophilicity, Surface Analysis, Coating, Contact Angle, Surface Tension, Acid-Base Interaction, Wettability, Dewatering Kinetics and Modeling.

ADVANCED CHEMICAL-MECHANICAL DEWATERING OF FINE PARTICLES

By

Ramazan Asmatulu

ABSTRACT

In the present work, novel dewatering aids and a novel centrifuge configuration were developed and applied for the purpose of dewatering fine particles. Three different types dewatering reagents were tested in different filtration and centrifugation units. These chemicals included low-HLB surfactants, naturally occurring lipids, and modified lipids. Most of these reagents are insoluble in water; therefore, they were used in solutions of appropriate solvents, such as light hydrocarbon oils and short-chain alcohols. The role of these reagents was to increase the hydrophobicity of the coal and selected mineral particles (chalcopryrite, sphalerite, galena, talc, clay, phosphate, PCC and silica) for the dewatering. In the presence of these reagents, the water contact angles on the coal samples were increased up to 90°. According to the Laplace equation, an increase in contact angle with the surfactant addition should decrease the capillary pressure in a filter cake, which should in turn increase the rate of dewatering and help reduce the cake moisture. The use of the novel dewatering aids causes a decrease in the surface tension of water and an increase in the porosity of the cake, both of which also contribute to improved dewatering.

A series of batch-scale dewatering tests were conducted on a variety of the coal and mineral samples using the novel dewatering aids. The results obtained with a Buchner funnel and air pressure filters showed that cake moistures could be reduced substantially, the extent of which depends on the particle size, cake thickness, drying time, reagent dosage, conditioning time, reagent type, sample aging, water chemistry, etc. It was determined that use of the novel dewatering aids could reduce the cake formation time by a significant degree due to the increased kinetics of dewatering. At the same time, the use of the dewatering aids reduced the cake moistures by allowing the water trapped in smaller capillaries of the filter cake. It was found that final cake moistures could be reduced by 50% of what can be normally achieved without using the reagents. However, the moisture reduction becomes difficult with increasing cake thickness. This problem can be minimized by applying a mechanical vibration to the cake,

spraying a short-chain alcohol on the cake and by adding a small amount of an appropriate coagulant, such as alum and CaCl_2 to the coal and mineral slurries.

The novel dewatering aids were also tested using several different continuous filters, including a drum filter, disc filter and horizontal belt filter (HBF). The results obtained with these continuous filtration devices were consistent with those obtained from the batch filters. Depending on the coal and mineral samples and the type of the reagent, 40 to 60% reductions in moisture were readily achieved.

When using vacuum disc filters, the cake thickness increased substantially in the presence of the novel dewatering aids, which could be attributed to the increased kinetics of dewatering. A dual vacuum system was developed in the present work in order to be able to control the cake thickness, which was necessary to achieve lower cake moistures. It was based on using a lower vacuum pressure during the cake formation time, while a full vacuum pressure was used during the drying cycle time. Thus, use of the dual vacuum system allowed the disc filter to be used in conjunction with the novel dewatering aids. Its performance was similar to that of HBF, which is designed to control cake thickness and cake formation time independently.

The effectiveness of using the novel dewatering aids were also tested in a full-continuous pilot plant, in which coal samples were cleaned by a flotation column before the flotation product was subjected to the disc filter. The tests were conducted with and without using novel dewatering aids. These results were consistent with those obtained from the laboratory and batch-scale tests.

The novel centrifuge developed in the present work was a unit, which combined a gravity force and air pressure. The new centrifuge was based on increasing the pressure drop across the filter cake formed on the surface of the medium (centrifuge wall). This provision made it possible to take advantage of Darcy's law and improve the removal of capillary water, which should help lower the cake moisture. A series of tests were conducted on several fine coal and mineral particles and obtained more than 50% moisture reduction even at very fine particle size ($2 \mu\text{m} \times 0$).

Based on the test results obtained in the present work, two proof-of-concept (POC) plants have been designed. The first was for the recovery of cyclone overflows that are currently being discarded in Virginia, and the other was for the recovery of fines from a pond in southern West Virginia. The former was designed based on the results of the plant tests conducted in the

present work. Cost vs. benefit analyses were conducted on the two POC plants. The results showed very favorable internal rates of return when using the novel dewatering aids.

Surface chemistry studies were conducted on the coal samples based on the results obtained in the present investigation. These consisted mainly of the surface characterization of the coal samples (surface mineral composition, surface area, zeta potential, x-ray photoelectron microscopy (XPS)), acid-base interactions of the solids and liquids, dewatering kinetic tests, contact angle measurements of the coal samples and surface force measurements using AFM. In addition, carbon coating on a silica plate using pulsed laser deposition (PLD) and Langmuir-Blodgett (LB) film deposition tests were conducted on the sample to better understand the surfactant adsorption and dewatering processes. The test results showed that the moisture reductions on the fine particles agree well with the surface chemistry results.

ACKNOWLEDGMENTS

The author would like to express his sincere appreciation to his advisor, Dr. Roe-Hoan Yoon, for the guidance, inspiration, suggestions, critics and financial support received through the end of this investigation. Without whose supports, this work would not have been possible. Special thanks are also given to his commute members, Dr. Luttrell, Dr. Adel, Dr. Weightman and Dr. Reynolds for their continued interests, useful suggestions, constructive comments and discussion to understand the overall process.

The author would like to acknowledge Drs. Cemal Biron, Bedri Ipekoglu, Guven Onal, Ertugrul Topuz, Cahit Coruh, Don McKeon, William Ducker, Mr. Bob Gentile, Mr. Tony Walters and Mrs. Shelley Nester for their kind help and guidance. He wishes to thank his colleagues Mr. Ismail Yildirim, Neeraj K. Mendiratta and Rajesh Pazhianur for teaching him the use of several equipments in the department, and the rest of his fellow students and employees at Holden Hall and Plantation Road for their friendship. Gratitude is expressed to my friends Aziz Dursun, Ilhan Tuzcu, Ferman Konukman, Yakup G. Gurbuz, Aysen Tulpar, Emre Isin and other TSA members for their friendless and continuous support. Appreciation is also extended to my dear house owner Mrs. Dorothy Foreman and her family for sharing their house and great assistance.

The financial support of this investigation from U.S. Department of Energy (Contract No: DE-AC26-98FT40153) and the Department of Mining and Minerals Engineering is sincerely appreciated.

Finally, the author would also like to express his most sincere appreciation to his mother Ayse, father Bayram for their years of hard work, financial supports and encouragement and sisters and brothers.

TABLE OF CONTENTS

	<u>Page</u>
TITLE PAGE.....	i
ABSTRACT.....	ii
ACKNOWLEDGMENTS.....	v
TABLE OF CONTENTS.....	vi
LIST OF FIGURES.....	x
LIST OF TABLES.....	xiii
CHAPTER 1 INTRODUCTION.....	1
1.1 General Background.....	1
1.2 Literature Review.....	3
1.2.1 General Information.....	3
1.2.2 Dewatering Theory.....	13
1.2.2.1 Filtration Theory.....	13
1.2.2.2 Centrifugation Theory.....	18
1.2.3 Surface Chemistry.....	20
1.2.3.1 Acid-Base Interaction In filtration.....	20
1.2.3.2 Hydrophobic forces in filtration.....	28
1.3 Solid-Liquid Separation Techniques.....	34
1.3.1 Sedimentation Techniques.....	34
1.3.2 Filtration Techniques.....	34
1.3.3 Centrifugation Techniques.....	36
1.3.4 Thermal Drying.....	38
1.4 Research Objectives.....	38
1.5 Report Organization	39
References.....	41
CHAPTER 2 FILTRATION	
2.1 Introduction.....	47
2.2 Experimental.....	54

2.2.1 Sample.....	54
2.2.2 Methods and Procedure.....	56
2.3 Results and Discussion.....	59
2.3.1 Dewatering Tests of Coal Samples.....	59
2.3.1.1 Low HLB Surfactant (Pure Surfactant).....	59
2.3.1.2 Natural Product.....	70
2.3.1.3 Modified Natural Product.....	82
2.3.1.4 Comparison with Conventional Dewatering Aids.....	85
2.3.2 Dewatering Tests of Mineral Samples.....	88
2.4 Summary and Conclusion.....	95
References.....	99
CHAPTER 3 CENTRIFUGATION.....	104
3.1 Introduction.....	104
3.2 Experimental.....	108
3.2.1 Samples.....	109
3.2.2 Equipment Design and Setup.....	111
3.3 Results and Discussion.....	111
3.3.1 Dewatering of Fine Coal Samples.....	111
3.3.2 Dewatering of Fine Mineral Samples.....	119
3.5 Summary and Conclusion.....	127
References.....	130
CHAPTER 4 ANALYSIS, EVALUATION AND PROCESSING.....	135
4.1 Introduction.....	135
4.2 Pilot Plant Configuration.....	135
4.3 Pilot Plan Tests.....	137
4.3.1 Experimental.....	137
4.3.1.1 Coal Sample.....	137
4.3.1.2 Method.....	138
4.3.2 Results and Discussions.....	143

4.3.2.1	Drum Filter.....	143
4.3.2.2	Horizontal Belt Filter.....	148
4.3.2.3	Disc Filter.....	154
4.4	Analysis, Evaluation and Processing.....	165
4.4.1	Sample Analysis.....	165
4.4.2	Process Evaluation.....	166
4.4.2.1	Technical Evaluation.....	167
4.4.2.2	Economical Evaluation.....	168
4.4.3	Conceptual Processing.....	172
4.4.3.1	Equipment Design.....	173
4.4.3.2	Circuit Design.....	176
4.4.3.3	Preliminary Cost Analysis.....	178
4.5	Summary and Conclusion.....	183
	References.....	185
 CHAPTER 5 SURFACE CHEMISTRY IN FILTRATION.....		189
5.1	Introduction.....	189
5.1.1	Background.....	189
5.1.2	Thermodynamics of Dewatering.....	192
5.1.3	Kinetics of Dewatering.....	196
5.2	Experimental.....	199
5.2.1	Material.....	199
5.2.2	Apparatus and Procedures.....	200
5.3	Results and Discussion.....	202
5.3.1	Characterization of the Coal Samples.....	202
5.3.2	Acid-Base Interactions on Solids and Liquids.....	210
5.3.3	Surface Coverage Studies.....	215
5.3.4	Surface Force Measurements.....	217
5.3.5	Dewatering Kinetics Tests.....	223
5.4	Summary and Conclusion.....	232
	References.....	235

CHAPTER 6 CONCLUSIONS	243
CHAPTER 7 FUTURE WORKS	248
LIST OF ACRONYMS AND ABBREVIATION.....	250
NOMENCLATURES.....	251
APPENDICES	252
APPENDIX A.....	252
APPENDIX B.....	261
APPENDIX C.....	266
APPENDIX D.....	269
APPENDIX E.....	271
APPENDIX F.....	274
APPENDIX G.....	275
APPENDIX H.....	288
APPENDIX I.....	289
APPENDIX J.....	294
VITA	297

LISTS OF FIGURES

Figure 1.1	Effects of particle size on the cost of dewatering coal	2
Figure 1.2	Schematic illustration of the electrical double layer	5
Figure 1.3	Schematic illustrations of surface charge due to the layer formations.....	6
Figure 1.4	Schematic representation of particle-water relationship in the cake structure.....	7
Figure 1.5	Relationship between pressure and moisture reduction.....	8
Figure 1.6	Formation of de-watering stages on the filter media.....	8
Figure 1.7	Schematic representation of the moisture zones between the particles	9
Figure 1.8	Typical results of dewatering tests, which show that cake moisture decreases rapidly at the beginning and then tapers off with increasing drying cycle time. The latter is due to the difficulty in removing the moisture from the capillaries of smaller radii. The situation gets worse with increasing cake thickness	15
Figure 1.9	Changes in the surface tension (γ) of water as a function of surfactant concentration in moles/liter.....	17
Figure 1.10	Changes in the water contact angles on the surfaces of various hydrophobicities as a function of the concentration of a) an anionic surfactant (sodium dodecylsulfate) and b) a cationic surfactant (dodecylammonium hydrochloride).....	18
Figure 1.11	Changes in pressure across the filter cake formed on a centrifugal filter.....	20
Figure 1.12	The schematic illustration of the solid/liquid/air interface.....	23
Figure 1.13	Hydrophobic Forces of coated silica plate and glass sphere at high contact angles.....	31
Figure 1.14	Effects of acid-base interaction on the surface free energy.....	33
Figure 1.15	Bird solid-bowl centrifuge used for solid-liquid separation.....	37
Figure 1.16	Bird screen-bowl centrifuge used for solid-liquid separation.....	38
Figure 2.1	Experimental setup for laboratory vacuum filtration tests.....	57
Figure 2.2	A schematic representation of the 3.5-inch diameter pressure filter.....	58
Figure 3.1	Schematic illustration of three acting pressures on the cake into the novel centrifuge	107

Figure 3.2	Schematic representation of the novel centrifuge with air pressure.....	110
Figure 3.3	Schematic representation of the novel centrifuge with air and vacuum pressures.....	111
Figure 4.1	The plant configuration used in the testing of dense medium products.....	136
Figure 4.2	The plant configuration used in the testing of classifying cyclone overflows.....	137
Figure 4.3	Photograph of the Sepor drum filte.....	139
Figure 4.4	Photograph of the Peterson disc filter.....	139
Figure 4.5	Figure illustrating the principle of the dual vacuum system.....	140
Figure 4.6	A schematic representation of a horizontal belt filter (HBF) with provisions for applying steam injection, reagent spray and mechanical vibration.....	141
Figure 4.7	Schematic of the test circuit used in the evaluation of the various filtration technologies.....	142
Figure 4.8	POC flowsheet for treating classifying cyclone overflow from existing preparation plants (Circuit I). (POC circuit enclosed within dashed box.).....	177
Figure 4.9	POC flowsheet for treating pond reclaim material (Circuit II).....	178
Figure 4.10	Effect of dewatering chemical cost on the economic indicators for Circuit II (pond reclaim).....	182
Figure 4.11	Effect of dewatering chemical cost on the economic indicators for Circuit I (plant retrofit).....	182
Figure 5.1	Effects of the capillary diameters on the formation of capillary forces.....	190
Figure 5.2	a) A schematic representation of dewatering process, b) equilibrium contact angle between three interfacial tensions.....	193
Figure 5.3	Schematic representation of kinetic test apparatus and test set-up.....	202
Figure 5.4	LORLM image analysis of a) Pittsburgh No: 8 coal sample and b) Elkview-Canada coal samples (0.4 x 0.3 mm).....	203
Figure 5.5	Schematic representation of XPS spectra (wide-scan) obtained for the Pittsburgh coal sample.....	205

Figure 5.6	Narrow-scan XPS spectra of the carbon compounds on the Pittsburgh coal sample.....	206
Figure 5.7	Narrow-scan XPS spectra of the sulfur compounds on the Pittsburgh coal sample.....	206
Figure 5.8	Schematic representation of XPS spectra (wide-scan) obtained for the Middle Fork coal sample.....	207
Figure 5.9	Curve fit C(1s) of the narrow-scan XPS spectra of the Middle Fork coal sample.....	208
Figure 5.10	Zeta potential of various untreated and treated a) West Virginia and b) Pittsburgh coal samples as a function of aluminum sulfate and Ph.....	209
Figure 5.11.	Results of the AFM image analysis of a) clean silica surface and b) PLD carbon coated silica surface.....	216
Figure 1.12	a) CaCl ₂ effects on the zeta potential of the silica particles as a function of pH, b) F/R vs H curves obtained between silica plate and glass sphere in the presence of CaCl ₂ at pH 9.75. The dashed lines represent the classical DLVO theory, and the symbols represent the experimental results fitted for the force data.....	218
Figure 5.13	Results of the AFM force measurements conducted between the glass sphere and silica plates in the presence of the reagent DAHCl at pH 5.8. The solid lines represent the classical DLVO theory, and the symbols represent the experimental results fitted for the force data.....	220
Figure 5.14	Results of the AFM force measurements conducted between the glass sphere and silica plates in the presence of the DAHCl and PMHO at pH 5.8 and ~40 °C. The solid lines represent the classical DLVO theory, and the symbols represent the experimental results fitted for the force data.....	221
Figure 5.15	Results of the AFM force measurements conducted between the glass sphere and silica plates in the presence of the DAHCl and Span 80 (dissolved 20% in diesel) at pH 5.8 and ~40 °C. The solid lines represent the classical DLVO theory, and the symbols represent the experimental results fitted for the force data.....	222
Figure 5.16	Effects of reagent Span 80 addition on the cake parameters of the coal sample at 60 kPa air pressure.....	224

LISTS OF TABLES

Table 1.1	Commonly Used Solid-liquid separation chemicals in dewatering processes.....	10
Table 1.2	Surface tension components of the selected liquids at 20 °C in mN/m.....	27
Table 1.3	Surface tension components of the selected solids at 20 °C in mN/m.....	28
Table 1.4	Surface Force Parameters on the symmetric interactions as a function of θ	32
Table 2.1	Selected Coal Samples Evaluated in the Dewatering Test Program.....	55
Table 2.2	Effects of Using Sorbitan Monooleate with Various Solvents for the Vacuum Filtration of a Pittsburgh Coal (0.5 mm x 0) Sample.....	59
Table 2.3	Effects of Using Ethyl Oleate for the Filtration of a 0.21 mm x 0 Elkview Coal Sample at 200 kPa of Air Pressure.....	60
Table 2.4	Effects of Using TDDP on the Filtration of an Elkview Coal at 200 kPa Air Press.....	61
Table 2.5	Effects of TDDP on the Surface Chemistry Parameters for the Filtration of a Pittsburgh Coal Sample.....	62
Table 2.6	Effects of Using Sorbitan Monooleate for the Filtration of a Blackwater Coal (0.6 mm x 0) Sample at Different Air Pressures.....	63
Table 2.7	Effects of Ultrasonic Vibration on the Vacuum Filtration of a Bituminous Coal (0.6 mm x 0) Using Sorbitan Monooleate on Massey Coal	64
Table 2.8	Effects of Spraying Different Reagents Over Middlefork Filter Cake When Using Sorbitan Monooleate and TDDP as Dewatering Aids.....	66
Table 2.9	Effects of Using TDDP Reagent, Butanol Spray, Vibration, and a Combination Thereof at 1.2-inch Cakes Thickness on a Massey Coal	67
Table 2.10	Effects of Using Electrolytes for the Filtration of Massey Bituminous Coal (0.2 mm x 0).....	68
Table 2.11	Effects of Using Electrolyte, Regent Spray, and Vibration on the Filtration of a Moss 3 Coal (0.6 mm x 0) at 1-inch Cake Thickness Using Sorbitan Monooleate as a Dewatering Aid.....	69
Table 2.12	Effects of Using Soybean Oil as a Dewatering Aid in Various Solvents on the Vacuum Filtration of a Pittsburgh Coal Sample (0.5 mm x 0)	71

Table 2.13	Synergistic Effect of Using Soybean Oil and Diesel Oil for the Vacuum Filtration of a Blackwater Coal Sample (0.85 mm x 0).....	72
Table 2.14	Effects of Using Different Vegetable Oils for the Vacuum Filtration of Elkview Coal.....	73
Table 2.15	Effects of Kerosene and Fish Oil on the Contact Angle of a Pittsburgh Coal Sample, Filtrate Surface Tension, and the Final Cake Moisture.....	74
Table 2.16	Effects of Using Fish Oil on the Filtration of a Coal Sample (0.6 mm x 0) from the Peak Downs Mine, Australia, at 200 kPa of Air Pressure	75
Table 2.17	Effects of Using Coconut Oil for the Filtration of a Bituminous Coal at 200 kPa of air Pressure.....	76
Table 2.18	Effects of Using Lard Oil for the Filtration of a 0.85 mm x 0 Massy Coal Sample at 100 and 200 kPa of Air Pressures.....	76
Table 2.19	Effects of Using a Sunflower Oil-Sorbitan Monooleate Blend for the Filtration of a Bituminous Coal Sample at 150 kPa Air Pressure	77
Table 2.20	Results of the Pressure Filtration Tests Conducted on the Elkview Flotation Product Using Soybean Oil.....	78
Table 2.21	Effects of Using Fish Oil on the Filtration of a Coal Sample (0.6 mm x 0) from the Peak Downs Mine, Australia, at 200 kPa of Air Pressure	78
Table 2.22	Effects of Vibration on the Filtration of a Virginia Coal Using Sunflower Oil as a Dewatering Aid	79
Table 2.23	Effects of Combining the Techniques of Using Sunflower Oil and Butanol Spray to Achieve Deep Moisture Reductions at 0.45-inch Cake Thickness.	80
Table 2.24	Effects of Combining the Techniques of Using Sunflower Oil, Al ³⁺ Ions, Butanol Spray, and Vibration to Achieve Deep Moisture Reductions at a 0.67-inch Cake Thickness.....	82
Table 2.25	Results of the Filtration Tests Conducted on a Meadow River Coal Sample (0.5 mm x 0) Using Sunflower Oil with and without Modification.....	83
Table 26	Results of the Vacuum Filtration Tests Conducted on a 0.5 mm x 0 Middle Fork Coal Sample Using a Modified Safflower Oil.....	84
Table 2.27	Results of the Filtration Tests Conducted on a 0.5 mm x 0 Meadow River Coal Sample Using a Modified Lard Oil	85

Table 2.28	Effects of Aging a Red River Coal Sample (0.6 mm x 0) for Using Span 80 and diesel.....	86
Table 2.29	Effects of Aging a Red River Coal Sample (0.6 mm x 0) for Using Superfloc 16.....	87
Table 2.30	Comparison the Buchner Filter Tests Conducted on the Middle Fork Coal Sample (0.6 mm x 0) Using High- and Low-HLB Surfactants as Filter Aids.....	87
Table 2.31	Effects of Using Fish Oil as a Dewatering Aid for the Filtration of a Zinc Concentrate from Boliden at 100 kPa of Air Pressure.....	89
Table 2.32	Effects of Using Sorbitan Monooleate Mixed with Diesel Oil for the Filtration of a Zinc Concentrate (0.105 mm x 0) at 100 kPa of Air Pressure and at Varying Reagent Addition and Cake Thickness.....	90
Table 2.33	Effects of Using Soybean Oil as a Dewatering Aid for the Vacuum Filtration of a Copper Concentrate.....	91
Table 2.34	Effects of Using Ethyl Oleate Mixed with Diesel Oil as a Dewatering Aid for the Vacuum Filtration of a Lead Concentrate (0.074 mm x 0) Sample at Varying Reagent Dosage and Cake Thickness.....	91
Table 2.35	Effects of Using Sun Flower Oil for the Filtration of a Talc (0.15 mm x 0) Sample at 200 kPa of Air Pressure with Varying Reagent Addition and Cake Thickness.....	92
Table 2.36	Flotation effects on silica dewatering conducted on a silica flour (0.038 mm x 0) using Span 80 dissolved 33.3% in diesel at 25 in.Hg vacuum pressure	93
Table 2.37	Effect of Span 80 (dissolved 33.3% in diesel) on the dewatering of the Brazilian Kaolin Clay (-2 micron) at 25 in.Hg Vacuum Pressure and 3 min Drying cycle time.....	94
Table 3.1	Effects of Using a Blend of TDDP and Soybean Oil as a Dewatering Aid for the Centrifugal Filtration of a Deslimed Bituminous Coal Sample at 2000 G.....	112
Table 3.2	Centrifugal test results of West Virginia coal sample using TDDP, Air Pressures at 2000 G-force.....	113

Table 3.3	Effects of Using a Dewatering Aid on the Centrifugal Filtration of Pittsburgh Coal at Different Air Pressures	114
Table 3.4	Effect of Using Compressed Air for the Centrifugal Filtration of a Pittsburgh Coal ³	115
Table 3.5	The Results Obtained with a Deslimed Microcel™ Flotation Product at 2,500 G and Varying Air Pressures.....	116
Table 3.6	Results Obtained with a Fine (-0.074 mm) Pittsburgh Coal Sample at 2,000 G and 0.5-inch Cake Thickness.....	117
Table 3.7	Effects of Centrifugal Forces on the Dewatering of a 1 mm x 0 Filter Feed at 300 k Pa of Air Pressure	118
Table 3.8	Effects of Spin Time for the Dewatering of a Pittsburgh Coal Sample at Different Air Pressures and at 2000 G.....	119
Table 3.9	The Results Obtained with a Sphalerite Concentrate at 2000 G and 0.62-inch Cake Thickness.....	120
Table 3.10	The Results Obtained with a Chalcopyrite Concentrate at 2000 G and 0.7-inches Cake Thickness.....	121
Table 3.11	The Centrifugal Results Obtained with a Galena Concentrate at 1000 G and 0.6-inch Cake Thickness.....	121
Table 3.12	Results Obtained on an East Georgia Kaolin Clay at 2000 G and 0.4-inch Cake Thickness.....	122
Table 3.13	The Results Obtained on a PCC Sample at 2000 G and 0.35-inch Cake Thickness	123
Table 3.14	Comparison of Using Vacuum and Air Pressures on the Centrifugal Filtration of a Phosphate Sample at 2000 G.....	124
Table 3.15	Results Obtained on a Phosphate Concentrate Using Both Compress Air and Vacuum Pressure at 2000G ¹	125
Table 3.16	Novel Centrifugal Test Results Conducted on Luzenac Clean Talc Sample* at 1500 G-Force.....	125
Table 3.17	Synergistic Effects of Using Centrifugal Force and Compressed Air for the Dewatering of a Talc Sample ³	126

Table 3.18	Centrifugal Test Results Conducted on Tennessee Silica Sample Using Air Pressure at 1500 G-Force	127
Table 4.1	Comparison of the Batch and Continuous Vacuum Filtration Tests Conducted on the Middle Fork Coal Sample Using Sorbitanmonooleate.....	144
Table 4.2	Effects of Using Ethyleneglycol Monooleate (EGMO) for the Filtration of a 28 Mesh x 0 Red River Coal Sample at 20-inch Hg Vacuum Pressure.....	145
Table 4.3	Effects of Using Sorbitanmonooleate for the Filtration of a 28 Mesh x 0 Red River Coal Sample at 22 in Hg Vacuum Pressure.....	146
Table 4.4	Effects of Using TridecylDihydrogen Phosphate (TDDP) for the Filtration of a 28 Mesh x 0 Red River Coal Sample at 22 in Hg Vacuum Pressure	146
Table 4.5	Effects of Using TDDP for the Filtration of a 0.5 mm x 0 Moss III Coal Sample in Conjunction with 20 g/ton Al ³⁺ Ions and Butanol Spray.....	147
Table 4.6	Effects of Using Oleic Acid for the Filtration of a 0.5 mm x 0 Middle Fork Coal Sample with Alum ¹ and Butanol Spray ²	148
Table 4.7	Results of the HBF Tests Conducted on the Elkview Coal Sample Using EGMO at 18- to 22-Inch Hg Vacuum Pressure.....	149
Table 4.8	Effect of Drying Cycle Time on the Use of Sorbitanmonooleate for the Dewatering of the Elkview Coal Sample Using Pilot-Scale Horizontal Belt Filter.....	150
Table 4.9	Effect of Conditioning Time for the Use of TDDP as Dewatering Aid in the Pilot-Scale HBF Tests on the 28 Mesh x 0 Red River Coal Sample ¹	151
Table 4.10	Effects of Using EGMO for the Dewatering of the 28 Mesh x 0 Pittsburgh Coal Using a Pilot-Scale HBF.....	152
Table 4.11	Effects of Using Oleic Acid for the Dewatering of the 28 Mesh x 0 Pittsburgh Coal Using a Pilot-Scale HBF.....	152
Table 4.12	Effects of Using Span 80 for the Dewatering of the 28 Mesh x 0 Pittsburgh Coal Using a Pilot-Scale HBF.....	153
Table 4.13	Effects of Using EGMO for the Dewatering of the 28 Mesh x 0 Pittsburgh Coal Using a Pilot-Scale HBF.....	153
Table 4.14	Effects of Using Reagent Esterified Lard Oil for the Dewatering of the 28 Mesh x	

	0 Pittsburgh Coal Using a Pilot-Scale HBF.....	154
Table 4.15	Results of the Pilot-Scale Filtration Tests Conducted on the DMS Product Pulverized in a Ball Mill to -28 Mesh Using Span 80.....	156
Table 4.16	Effects of Using Span 80 for the Pilot-Scale Vacuum Disc Filter Tests Conducted on the 28 Mesh x 0 Pittsburgh Coal	157
Table 4.17	Effects of Using TDDP for the Pilot-Scale Vacuum Disc Filter Tests Conducted on the 28 Mesh x 0 Pittsburgh Coal.....	157
Table 4.18	Effects of Using EGMO for the Pilot-Scale Vacuum Disc Filter Tests Conducted on the 28 Mesh x 0 Pittsburgh Coal.....	158
Table 4.19	Effects of Using EGMO for the Pilot-Scale Vacuum Disc Filter Tests Conducted on the 28 Mesh x 0 Pittsburgh Coal.....	159
Table 4.20	Effect of of Drying Cycle Time on the Dewatering of the 28 Mesh x 0 Pittsburgh Coal Sample Using Oleic Acid as Dewatering Aid.....	160
Table 4.21	Effect of %Solids in the Feed to a 24-Inch Diameter Vacuum Disc Filter When Using Oleic Acid as Dewatering Aid.....	161
Table 4.22	Effect of Using Span 80 for the Dewatering of Elkview Coal Using a Pilot-Scale Vacuum Disc Filter.....	162
Table 4.23	Pilot-Scale Disc Filter Tests Conducted on the 28 Mesh x 0 Red River Coal Sample Using Span 80.....	162
Table 4.24	Pilot-Scale Disc Filter Tests Conducted on the 28 Mesh x 0 Mesh Red River Coal Sample Using EGMO.....	163
Table 4.25	Disc filter test result on dewatering of Moss 3 coal sample* by using Span 80 (33.3% in diesel) at 24/19 in Hg Vacuum Pressure.....	164
Table 4.26	Pilot-Scale Disc Filter Tests Conducted on the 28 Mesh x 0 Pittsburgh Coal Sample Using Esterified Lard Oil.....	164
Table 4.27	Comparison of product yields obtained using the Buchaner and air pressure filters.....	165
Table 4.28	Effect of Improved Dewatering on the overall Plant Yield and Profitability 10.0% Reduction in Moisture.....	168
Table 4.29	Effect of Improved Dewatering on the Overall Plant Yield and Profitability 15% Reduction in Moisture.....	170

Table 4.30	Capital Cost Estimate for the Retrofit of an Existing Plant to Clean and Dewater - 100 mesh raw coal.....	171
Table 4.31	Annual Operating and Maintenance Costs.....	172
Table 4.32	Financial Analysis.....	173
Table 4.33	Scale-Up Calculations for the POC Column Flotation Installation.....	174
Table 4.34	Scale-Up Calculations for the POC Disc Filter Installation.....	175
Table 4.35	Cost-Benefit Analysis for the POC Circuits.....	181
Table 5.1	HLB numbers of hydrophilic and lipophilic groups,.....	191
Table 5.2	HLB numbers of selected commercial surfactants.....	191
Table 5.3	Effects of time on the coal samples in water.....	203
Table 5.4	Effects of surface area on dewatering of Pittsburgh and West Virginia coal sample using Span 80 at 100 kPa pressure.....	204
Table 5.5	Atomic concentration of the coal samples obtained by XPS analyses.....	208
Table 5.6	Effects of Using Aluminum Sulfate for the Filtration of West Virginia and Pittsburgh Coal Sample.....	210
Table 5.7	The values of the contact angles on the coal samples* using known liquids.....	211
Table 5.8	The surface tension components of different region coal samples.....	212
Table 5.19	The surface tension components of reagents used as dewatering aids.....	213
Table 5.10	The surface tension components of the known solids and liquids.....	214
Table 5.11	The effects of acid-base interactions on dewatering of Middle Fork coal sample using different reagents at 100 kPa air pressure.....	214
Table 5.12	Dewatering test results on silica sample* using CaCl ₂ and CaO at 25 inHg Vacuum Pressure and pH 9.75.....	219
Table 5.13	Dewatering test results on silica sample* using dodecylamine hydrochloride and dewatering aids at 25 in. Hg vacuum pressure and pH 5.8.....	222
Table 5.14	Effect of reagent dosages on the cake parameters of Pittsburgh coal sample using Reagent Span 80 (dissolved 33.3% in diesel) at 60 kPa air pressure.....	225
Table 5.15	Effect of reagent dosages on the cake parameters of Elkview coal sample using Reagent TDDP (dissolved 33.3% in diesel) at 120 kPa air pressure.....	230

Table 5.16	Effect of reagent dosages on the cake parameters of West Virginia coal sample using Reagent ROE (dissolved 33.3% in diesel) at 80 kPa air pressure.....	232
------------	---	-----